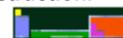


**Revised August 2012**



## AP LAB 4b: Finding the formula of a hydrated salt



**Aim** To determine the formula of a hydrated salt

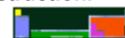
**Apparatus** Test tube, test tube holder, Bunsen burner, spatula, goggles, electronic balance

**Chemicals** Hydrated copper(II) sulfate

### **Method**

1. Weigh an empty test tube on the electronic balance. Record the mass.
2. Add approximately 0.30 g of the hydrated copper(II) sulfate to the test tube and record the total mass.
3. To avoid splattering, *gently* heat the test tube and its contents using the blue Bunsen flame (by wafting the test tube back and forth in the flame) for approximately three minutes, record your observations.
4. **Allow the test tube to cool for about five minutes.** Record the mass using the electronic balance. **DO NOT PLACE HOT TEST TUBES ON THE ELECTRONIC BALANCES.**
5. Repeat steps #3 and #4 as necessary, until a constant mass is obtained. (A constant mass is one that after repeated, heating, cooling and re-weighing, stays the same).

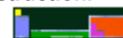
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### Results

Observations

Mass of empty test tube	
Mass of test tube + hydrated salt BEFORE heating	
Mass of test tube + salt after first heating and cooling cycle	
Mass of test tube + salt after second heating and cooling cycle	
Mass of test tube + salt after third heating and cooling cycle	
Mass of test tube + salt after fourth heating and cooling cycle	
Mass of test tube + salt after fifth heating and cooling cycle	
Mass of test tube + salt after sixth heating and cooling cycle.	
etc.	



**Conclusion/Calculation**

1. Perform the calculations below.

<b>(i) Calculate the mass of hydrated salt added to the test tube at the start of the experiment.</b>
<b>(ii) Calculate the mass of the anhydrous (without water) salt at the end of the experiment.</b>
<b>(iii) Calculate the mass of water driven off by heating the hydrated salt.</b>
<b>(iv) Calculate the moles of the anhydrous (without water) salt present.</b>
<b>(v) Calculate the moles of water driven off by heating.</b>
<b>(vi) Compare your answers to (iv) and (v) above in order to determine the formula of the hydrated salt.</b>

2. In another similar experiment, a student heats the hydrated salt carelessly and some of the salt spits out of the test tube. What would be the effect of this experimental error on the calculation of the formula of the hydrated salt?