

REGULAR LAB 1b: Density (Significant Figures)

Aim To calculate the density of three liquids

Apparatus 10 mL graduated cylinder, electronic balance, pipets

Chemicals Water, rubbing alcohol (propan-2-ol) and vinegar (ethanoic acid)

Method

1. Fill a new pipet with water, dry the outside of the pipet and record the mass of the entire unit.
2. Empty the water into the graduated cylinder and record the volume. (Record the volume to only one decimal place).
3. Record the mass of the empty pipet (Record the mass to two decimal places).
4. Calculate the density (in g/mL) of water, observing the rules for significant figures.
5. Repeat steps #1 through #4 twice. Then determine the average density of water, again observing the rules for significant figures.
6. Repeat steps #1 through #5 using rubbing alcohol instead of water, and then again using vinegar instead of water.

Results

	1 st trial	2 nd trial	3 rd trial
Mass of full pipet in g			
Mass of empty pipet in g			
Mass of water in g			
Volume of water in mL			
Density of water in g/mL			

Average density of water in g/mL = _____

	1 st trial	2 nd trial	3 rd trial
Mass of full pipet in g			
Mass of empty pipet in g			
Mass of alcohol in g			
Volume of alcohol in mL			
Density of alcohol in g/mL			

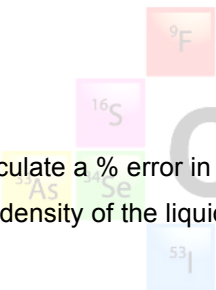
Average density of rubbing alcohol in g/mL = _____

	1 st trial	2 nd trial	3 rd trial
Mass of full pipet in g			
Mass of empty pipet in g			
Mass of acid in g			
Volume of acid in mL			
Density of acid in g/mL			

Average density of vinegar in g/mL = _____

Conclusion/Calculation

1. What is the purpose of repeating each experiment three times?
2. Why is it important to dry the outside of the pipet each time before massing it?
3. Calculate a % error in each experiment, by using a reliable reference source for the actual value of the density of the liquids used.



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