

## AP WORKSHEET 00D: ANSWERS

1. $K_2Cr_2O_7$
2. $Li_2S$
3. $KBr$
4. $CsI$
5. $Ca_3P_2$
6. $NaF$
7. $SrO$
8. $BeS$
9. $MgBr_2$
10. $Li_2O$
11. $SrCl_2$
12. $BaBr_2$
13. $MgS$
14. $MgI_2$
15. $HF$
16. $Ba_3P_2$
17. $Na_2HPO_4$
18. $KCl$
19. $Li_3N$
20. $CaS$
21. $Rb_2O$
22. $Sr_3N_2$
23. $Cs_3P$
24. $MgCO_3$
25. $BeSO_4$

26. $\text{N}_2\text{O}_4$
27. $\text{CO}_2$
28. $\text{Hg}_2\text{Cl}_2$
29. $\text{HI}$
30. $\text{HIO}_3$
31. $\text{HBrO}_4$
32. $\text{HBrO}$
33. $\text{PCl}_5$
34. $\text{ICl}$
35. $\text{SbF}_3$
36. $\text{BrF}$
37. $\text{BrO}_2$
38. $\text{N}_2\text{O}_5$
39. $\text{CS}$
40. $\text{TeO}_2$
41. $\text{PBr}_3$
42. $\text{Cl}_4$
43. $\text{V}_2(\text{CrO}_4)_5$
44. $\text{ZnCO}_3$
45. $\text{AgOH}$
46. $\text{V}_2(\text{CrO}_4)_3$
47. $\text{HgI}_2$
48. $\text{U}(\text{NO}_3)_5$
49. $\text{NiN}$
50. $\text{H}_2\text{SO}_4$

ADRIAN DINGLE'S  
Chemistry Pages

51. Scandium (III) chloride
52. Hydrochloric acid, Hydrogen chloride, Hydrogen monochloride <sub>1</sub>
53. Platinum (IV) oxide
54. Antimony (V) chlorate <sub>4</sub>
55. Germanium disulfide
56. Zinc oxide <sub>2</sub>
57. Vanadium (II) sulfate
58. Copper (II) chloride
59. Titanium (IV) oxide
60. Nickel (III) nitride
61. Nickel (II) phosphate
62. Cobalt (III) fluoride
63. Gold (III) oxide
64. Zinc phosphide <sub>2</sub>
65. Chromium (VI) nitrate
66. Sodium iodite
67. Sodium iodate
68. Sodium iodide
69. Sulfurous acid
70. Carbonic acid
71. Aluminum nitride
72. Aluminum hydride
73. Lithium arsenate
74. Sodium cyanide
75. Sodium peroxide

76. Lithium phosphite
77. Potassium hydrogen carbonate
78. Hydrofluoric acid, Hydrogen fluoride, Hydrogen monofluoride <sub>1</sub>
79. Gold (II) iodide
80. Potassium permanganate
81. Sodium dichromate
82. Silver chromate <sub>2</sub>
83. Silver chloride <sub>2</sub>
84. Sodium ethanoate
85. Radium fluoride
86. Potassium thiocyanate
87. Iron (II) sulfide
88. Iron (III) sulfite
89. Iron (II) sulfate
90. Magnesium sulfide
91. Sodium thiosulfate
92. Rubidium chloride
93. Copper (II) hydroxide
94. Magnesium nitride
95. Copper (I) nitride
96. Lithium hydride
97. Potassium oxide
98. Potassium peroxide
99. Lithium nitride
100. Darmstadtium (III) chloride

Notes:

1. Strict nomenclature rules suggest mono should be applied to the second element in these compounds, but this is almost always dropped. Additionally these compounds can also be named as binary acids when they are in solution. This distinction is important in subsequent topics but less so here.
2. Although silver and zinc are d-block elements they do not commonly exhibit charges other than +1 and +2 respectively in compounds. As a result (I) and (II) are commonly omitted from nomenclature.
3. Here antimony is being treated as a metal hence the charge is stated, in other instances it may be treated as a non-metal. If so an alternative name here would be antimony trifluoride.
4. Here antimony is paired with the chlorate ion, so it is reasonable to assume it carries a charge too, hence the use of (V). See note 3.
5. Ds is the symbol for element 110 that was named in 2003. Here it is assumed it could take on a number of different charges in a compound, one of which is +3.