



4. Calculate the volume of 3.00 M sulfuric acid that must be diluted with water to produce 500. mL of 2.25 M sulfuric acid. (4)

5. Calculate the volume of water that must be added to 2.75 M NaOH in order to produce 150. mL of 2.00 M NaOH. (4)

6. Outline the correct laboratory procedure for diluting a stock solution of 18.4 M sulfuric acid to produce 1.0 L of 1.84 M sulfuric acid. Take care to include any calculations, safety procedures and how you would use appropriate glassware in your answer. (4)