

AP WORKSHEET 04B: Balancing Equations II

Balance the following equations using the **lowest possible integers**. (20)

1. $\text{N}_2 + \text{O}_2 \rightarrow \text{NO}$
2. $\text{C}_{10}\text{H}_{20} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
3. $\text{C}_4\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
4. $\text{Al}(\text{OH})_3 + \text{HCl} \rightarrow \text{AlCl}_3 + \text{H}_2\text{O}$
5. $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
6. $\text{Li} + \text{AlBr}_3 \rightarrow \text{LiBr} + \text{Al}$
7. $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
8. $\text{NH}_4\text{OH} + \text{H}_3\text{PO}_4 \rightarrow (\text{NH}_4)_3\text{PO}_4 + \text{H}_2\text{O}$
9. $\text{Li} + \text{P} \rightarrow \text{Li}_3\text{P}$
10. $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
11. $\text{Al}(\text{OH})_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
12. $\text{K} + \text{Cl}_2 \rightarrow \text{KCl}$
13. $\text{Li} + \text{S}_8 \rightarrow \text{Li}_2\text{S}$
14. $\text{H}_3\text{PO}_4 + \text{Mg}(\text{OH})_2 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
15. $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
16. $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
17. $\text{Ca}_3(\text{PO}_4)_2 + \text{SiO}_2 + \text{C} \rightarrow \text{CaSiO}_3 + \text{CO} + \text{P}$
18. $\text{NH}_3 + \text{O}_2 \rightarrow \text{NO} + \text{H}_2\text{O}$
19. $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
20. $\text{C} + \text{SO}_2 \rightarrow \text{CS}_2 + \text{CO}$